

PRACTICE

Answers in Appendix E

1. Determine the pH of the following solutions:
 - a. 1×10^{-3} M HCl
 - b. 1.0×10^{-5} M HNO₃
 - c. 1×10^{-4} M NaOH
 - d. 1.0×10^{-2} M KOH

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Answers in Appendix E

1. What is the pH of a solution if the $[\text{H}_3\text{O}^+]$ is $6.7 \times 10^{-4} \text{ M}$?
2. What is the pH of a solution with a hydronium ion concentration of $2.5 \times 10^{-2} \text{ M}$?
3. Determine the pH of a $2.5 \times 10^{-6} \text{ M}$ HNO_3 solution.
4. Determine the pH of a $2.0 \times 10^{-2} \text{ M}$ $\text{Sr}(\text{OH})_2$ solution.