

given in Table 3.

## PRACTICE

*Answers in Appendix E*

1. Calculate the solubility product constant,  $K_{sp}$ , of lead(II) chloride,  $\text{PbCl}_2$ , which has a solubility of 1.0 g/100. g  $\text{H}_2\text{O}$  at  $20^\circ\text{C}$ .
2. A 5.0 gram sample of  $\text{Ag}_2\text{SO}_4$  will dissolve in 1.0 L of water. Calculate the solubility product constant for this salt.

**extend**

Go to [go.h](#)  
more practice  
that ask you  
solubility pro  
constants.



Keywords