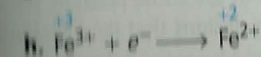
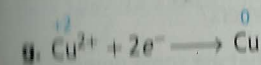
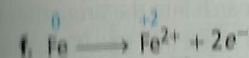
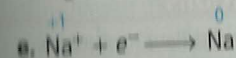
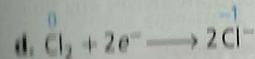
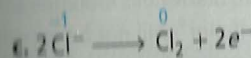
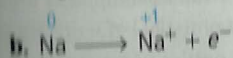
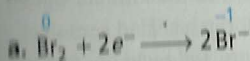
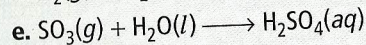
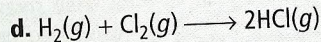
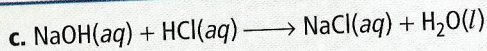
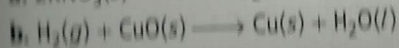
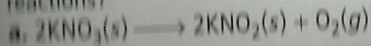


SECTION REVIEW

- How are oxidation numbers assigned?
- Label each of the following half-reactions as either an oxidation or a reduction half-reaction:



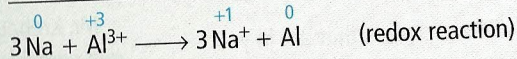
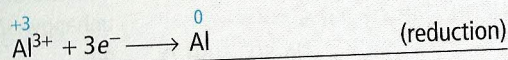
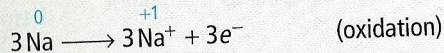
- Which of the following equations represent redox reactions?



- For each redox equation identified in the previous question, determine which element is oxidized and which is reduced.

Critical Thinking

- ANALYZING INFORMATION** Use the following equations for the redox reaction between aluminum metal and sodium metal to answer the questions below.



- Explain how this reaction illustrates that charge is conserved in a redox reaction.
- Explain how this reaction illustrates that mass is conserved in a redox reaction.
- Explain why electrons do not appear as reactants or products in the combined equation.